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**CHEMISTRY
HIGHER LEVEL
PAPER 1**

Thursday 8 May 2008 (afternoon)

1 hour

INSTRUCTIONS TO CANDIDATES

- Do not open this examination paper until instructed to do so.
- Answer all the questions.
- For each question, choose the answer you consider to be the best and indicate your choice on the answer sheet provided.
- The periodic table is provided for reference on page 2 of this examination paper.

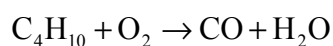
The Periodic Table

1	2	3	4	5	6	7	0	
<div>1 H 1.01</div>		<div>Atomic Number</div> <div>Element</div> <div>Atomic Mass</div>					<div>2 He 4.00</div>	
<div>3 Li 6.94</div>	<div>4 Be 9.01</div>						<div>9 F 19.00</div>	
<div>11 Na 22.99</div>	<div>12 Mg 24.31</div>						<div>10 Ne 20.18</div>	
<div>19 K 39.10</div>	<div>20 Ca 40.08</div>	<div>21 Sc 44.96</div>	<div>22 Ti 47.90</div>	<div>23 V 50.94</div>	<div>24 Cr 52.00</div>	<div>25 Mn 54.94</div>	<div>18 Ar 39.95</div>	
<div>37 Rb 85.47</div>	<div>38 Sr 87.62</div>	<div>39 Y 88.91</div>	<div>40 Zr 91.22</div>	<div>41 Nb 92.91</div>	<div>42 Mo 95.94</div>	<div>43 Tc 98.91</div>	<div>36 Kr 83.80</div>	
<div>55 Cs 132.91</div>	<div>56 Ba 137.34</div>	<div>57[†] La 138.91</div>	<div>72 Hf 178.49</div>	<div>73 Ta 180.95</div>	<div>74 W 183.85</div>	<div>75 Re 186.21</div>	<div>54 Xe 131.30</div>	
<div>87 Fr (223)</div>	<div>88 Ra (226)</div>	<div>89[‡] Ac (227)</div>						<div>86 Rn (222)</div>
[†]								
		<div>58 Ce 140.12</div>	<div>59 Pr 140.91</div>	<div>60 Nd 144.24</div>	<div>61 Pm 146.92</div>	<div>62 Sm 150.35</div>	<div>71 Lu 174.97</div>	
				<div>63 Eu 151.96</div>	<div>64 Gd 157.25</div>	<div>65 Tb 158.92</div>	<div>70 Yb 173.04</div>	
				<div>66 Dy 162.50</div>	<div>67 Ho 164.93</div>	<div>68 Er 167.26</div>	<div>69 Tm 168.93</div>	
[‡]								
		<div>90 Th 232.04</div>	<div>91 Pa 231.04</div>	<div>92 U 238.03</div>	<div>93 Np (237)</div>	<div>94 Pu (242)</div>	<div>103 Lr (260)</div>	
				<div>95 Am (243)</div>	<div>96 Cm (247)</div>	<div>97 Bk (247)</div>	<div>102 No (259)</div>	
				<div>98 Cf (251)</div>	<div>99 Es (254)</div>	<div>100 Fm (257)</div>	<div>101 Md (258)</div>	

1. How many molecules are present in a 9.0 g sample of water?

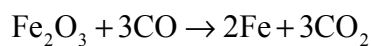
- A. 0.5
- B. 1.0
- C. 6.0×10^{23}
- D. 3.0×10^{23}

2. What is the coefficient for oxygen when this equation is balanced using the lowest whole number?



- A. 4
- B. 5
- C. 9
- D. 13

3. What is the maximum mass of iron that can be produced from the reduction of 80 tonnes of iron(III) oxide ($M_r = 160$), based on this equation?



- A. 28 tonnes
- B. 56 tonnes
- C. 84 tonnes
- D. 112 tonnes

4. Which species represent a pair of isotopes?

Species	Number of protons	Number of electrons	Number of neutrons
L	12	12	12
M	13	13	13
P	13	10	13
Q	12	12	14

- A. L and M
- B. L and P
- C. P and Q
- D. L and Q

5. How many unpaired electrons are there in the Co^{2+} ion?

- A. 7
- B. 5
- C. 3
- D. 2

6. Which processes occur in the mass spectrometer?

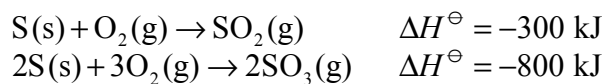
- I. Ionization by electron bombardment
 - II. Acceleration by a magnetic field
 - III. Deflection by a magnetic field
- A. I and II only
 - B. I and III only
 - C. II and III only
 - D. I, II and III

7. Which properties decrease in value when descending group 1?
- I. Atomic radius
 - II. Ionization energy
 - III. Electronegativity
- A. I and II only
 - B. I and III only
 - C. II and III only
 - D. I, II and III
8. The ionization energies of three consecutive elements in the periodic table are 1680, 2080 and 494 kJ mol⁻¹ respectively. Which of the following shows the elements with these values?
- A. O F Ne
 - B. F Ne Na
 - C. Ne Na Mg
 - D. Na Mg Al
9. Which comparison of radii of atoms and ions is correct?
- A. Cl⁻ > Cl
 - B. H⁺ > H⁻
 - C. Na⁺ > Na
 - D. Mg²⁺ > Mg⁺

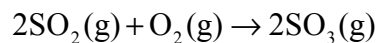
10. Which trend is correct when the elements are considered from left to right across period 3?
- The acidic character of the oxides decreases.
 - The electrical conductivity of the elements increases.
 - The bonding of the chlorides changes from ionic to covalent.
 - Electronegativity decreases.
11. Which substance will **not** conduct an electric current?
- C(s)(graphite)
 - NaF(l)
 - CaO(s)
 - KI(aq)
12. Which of the following liquids is non-polar?
- Water
 - Hexane
 - Propanone
 - Ethanol
13. The following substances all contain a nitrogen to nitrogen bond: N_2 , N_2H_4 , N_2H_2 . Which shows them in **increasing** order of nitrogen to nitrogen bond length (smallest first)?
- N_2H_4 , N_2H_2 , N_2
 - N_2 , N_2H_2 , N_2H_4
 - N_2H_2 , N_2H_4 , N_2
 - N_2H_4 , N_2 , N_2H_2

14. What is the bond angle in NO_3^- ?
- A. 107°
 - B. 109.5°
 - C. 120°
 - D. 180°
15. The temperature of 1 dm^3 of a gas is increased from 32°C to 64°C at constant pressure. What is the new volume in dm^3 ?
- A. 1.1
 - B. 1.3
 - C. 1.6
 - D. 2.0
16. Which change does **not** lead to an increase in entropy?
- A. Mixing nitrogen and oxygen gases at room temperature
 - B. Cooling steam so that it condenses to water
 - C. Heating hexane to its boiling point
 - D. Dissolving sugar in water

17. The enthalpy changes for two reactions are shown below.



What is the enthalpy change for this reaction in kJ?



- A. –200
 - B. –500
 - C. –1100
 - D. –1400
18. Which process is exothermic?
- A. $\text{Na(s)} \rightarrow \text{Na(g)}$
 - B. $\text{Ca(g)} \rightarrow \text{Ca}^+\text{(g)} + \text{e}^-$
 - C. $\text{Br(g)} + \text{e}^- \rightarrow \text{Br}^-\text{(g)}$
 - D. $\text{I}_2\text{(g)} \rightarrow 2\text{I(g)}$
19. Which equation represents the standard enthalpy of formation of calcium fluoride?
- A. $\text{Ca(g)} + \text{F}_2\text{(g)} \rightarrow \text{CaF}_2\text{(g)}$
 - B. $\text{Ca(s)} + \text{F}_2\text{(g)} \rightarrow \text{CaF}_2\text{(s)}$
 - C. $\text{Ca}^{2+}\text{(g)} + 2\text{F}^-\text{(g)} \rightarrow \text{CaF}_2\text{(s)}$
 - D. $\text{Ca}^{2+}\text{(s)} + 2\text{F}^-\text{(g)} \rightarrow \text{CaF}_2\text{(s)}$

20. 25 cm³ of 1.0 mol dm⁻³ NaOH is added to 25 cm³ of 1.0 mol dm⁻³ HCl. The temperature rise is 6°C. Which reactants will also give a temperature rise of 6°C?

- A. 25 cm³ of 2.0 mol dm⁻³ NaOH and 25 cm³ of 2.0 mol dm⁻³ HCl.
- B. 50 cm³ of 1.0 mol dm⁻³ NaOH and 50 cm³ of 1.0 mol dm⁻³ HCl.
- C. 50 cm³ of 0.5 mol dm⁻³ NaOH and 50 cm³ of 0.5 mol dm⁻³ HCl.
- D. 100 cm³ of 0.25 mol dm⁻³ NaOH and 100 cm³ of 0.25 mol dm⁻³ HCl.

21. Which reaction is the most exothermic?

- A. $\text{Li}^+(\text{g}) + \text{F}^-(\text{g}) \rightarrow \text{LiF}(\text{s})$
- B. $\text{Na}^+(\text{g}) + \text{Cl}^-(\text{g}) \rightarrow \text{NaCl}(\text{s})$
- C. $\text{Mg}^{2+}(\text{g}) + \text{O}^{2-}(\text{g}) \rightarrow \text{MgO}(\text{s})$
- D. $\text{Ca}^{2+}(\text{g}) + \text{S}^{2-}(\text{g}) \rightarrow \text{CaS}(\text{s})$

22. The table shows data for a reaction between X and Y.

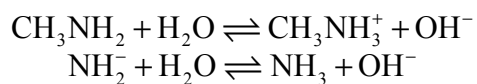
Experiment	[X] mol dm ⁻³	[Y] mol dm ⁻³	Rate of reaction mol dm ⁻³ s ⁻¹
1	0.4	0.24	1.2×10^{-4}
2	0.8	0.24	2.4×10^{-4}
3	0.4	0.12	3.0×10^{-5}

The overall order of reaction is:

- A. 1
- B. 2
- C. 3
- D. 4

23. Which units could be used for the rate of a chemical reaction?
- A. $\text{mol dm}^{-3} \text{ min}$
 - B. $\text{mol}^{-1} \text{ min}^{-1}$
 - C. $\text{dm}^3 \text{ min}$
 - D. $\text{mol dm}^{-3} \text{ min}^{-1}$
24. 10 cm^3 of liquid hexane is placed in a closed 1 dm^3 container at 298 K . Which change would increase the equilibrium vapour pressure of the hexane in the container?
- A. Putting the container in a refrigerator
 - B. Adding 10 cm^3 of hexane to the container
 - C. Reducing the volume of the container to 0.5 dm^3
 - D. Putting the container in a water bath at 308 K
25. Which change will increase the equilibrium concentration of sulfur trioxide in this reaction?
- $$2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g}) \quad \Delta H^\ominus = \text{negative}$$
- A. Decreasing the concentration of oxygen
 - B. Increasing the pressure
 - C. Using a catalyst
 - D. Increasing the temperature

26. Which species act as Brønsted – Lowry bases in the following reactions?



- I. CH_3NH_2
- II. CH_3NH_3^+
- III. NH_2^-

- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

27. The ionic product constant of water at 45°C is $4 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$. Which statement is correct about pure water at 45°C?

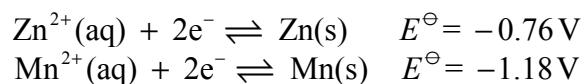
- A. $\text{pH} = 7$
- B. $[\text{H}^+] = [\text{OH}^-]$
- C. $[\text{OH}^-] > [\text{H}^+]$
- D. $[\text{H}^+] > [\text{OH}^-]$

28. A weak monoprotic acid is 10 % dissociated in a solution of concentration 0.01 mol dm^{-3} . What is the pH value of the solution?

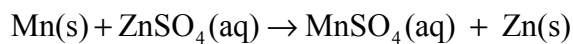
- A. 0.1
- B. 1.0
- C. 2.0
- D. 3.0

- 29.** Which change increases the pH of a solution from 3 to 6?
- A. Doubling the $[\text{H}^+]$
 - B. Halving the $[\text{OH}^-]$
 - C. Decreasing the $[\text{H}^+]$ by a factor of 1000
 - D. Decreasing the $[\text{OH}^-]$ by a factor of 1000
- 30.** Which pair of compounds, in aqueous solution, could be used to make a buffer solution?
- A. CH_3COOH and HCl
 - B. HCl and NaOH
 - C. HCl and NH_4Cl
 - D. HCOOH and NaOH
- 31.** During the electrolysis of aqueous sulfuric acid, 1g of hydrogen gas forms at the negative electrode. What mass in grams of oxygen forms at the positive electrode in the same time?
- A. 4
 - B. 8
 - C. 16
 - D. 32
- 32.** Which is the strongest oxidizing agent?
- A. I_2
 - B. I^-
 - C. F_2
 - D. F^-

33. The following are standard electrode potentials.



What is the E^{\ominus} for this reaction?



- A. -0.42 V
 B. $+0.42 \text{ V}$
 C. -1.94 V
 D. $+1.94 \text{ V}$
34. Which compound cannot be easily oxidized using acidified potassium dichromate(VI) solution?
- A. $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$
 B. $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$
 C. $(\text{CH}_3)_2\text{CHCH}_2\text{OH}$
 D. $(\text{CH}_3)_3\text{COH}$
35. In which reaction does hydrogen act as an oxidizing agent?
- A. $\text{Ca} + \text{H}_2 \rightarrow \text{CaH}_2$
 B. $\text{F}_2 + \text{H}_2 \rightarrow 2\text{HF}$
 C. $\text{C}_2\text{H}_2 + \text{H}_2 \rightarrow \text{C}_2\text{H}_4$
 D. $\text{O}_2 + 2\text{H}_2 \rightarrow 2\text{H}_2\text{O}$

36. Which species **cannot** act as a nucleophile?

- A. H_2O
- B. NH_3
- C. CN^-
- D. CH_4

37. Which compounds show three main peaks in their ^1H NMR spectra?

- I. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
- II. $\text{CH}_3\text{CH}_2\text{COOH}$
- III. $(\text{CH}_3)_3\text{CCH}_2\text{CH}_2\text{Br}$

- A. I and II only
- B. I and III only
- C. II and III only
- D. I, II and III

38. Which halogenoalkane reacts most rapidly with silver nitrate solution to form a precipitate?

- A. 1-bromobutane
- B. 1-iodobutane
- C. 2-bromo-2-methylpropane
- D. 2-iodo-2-methylpropane

39. Which is the correct formula of 2,3-dichloro-2-methylpentane?
- A. $\text{CH}_3\text{CCl}(\text{CH}_3)\text{CHClCH}_2\text{CH}_3$
 - B. $\text{CH}_3\text{CH}(\text{CH}_3)\text{CCl}_2\text{CH}_2\text{CH}_3$
 - C. $\text{CH}_3\text{CCl}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}_3$
 - D. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CHClCHClCH}_3$
40. What type of reaction occurs when hexanedioic acid and 1,6-diaminohexane react together to form nylon?
- A. Addition
 - B. Condensation
 - C. Esterification
 - D. Substitution
-